**MOUNT LITERA ZEE SCHOOL, HARIDWAR**

 **Session- 2018-19**

**Subject**- Science **Revision Worksheet-1 Date :**

**Topic**-Chemical reactions **Class-10** & equations

**Q-1 (i) What type of reaction is burning of magnesium with oxygen?**

 **(ii) Name the product formed when ferrous sulphaye is heated.**

**Q-2 A shiny brown colored element X on heating in air becomes black in color. Name the element X and the black colored compound formed.**

**Q-3 Write the decomposition reaction of silver chloride. What type of decomposition is this, also given the name of two salts of silver halides which are used in black & white photography?**

**Q-4 Why the blue color of copper sulphate fades by adding iron nails in it also write the chemical equation involved?**

**Q-5 What do you mean by double displacement reaction? What is the color of barium sulphate precipitate in the reaction of barium chloride with sodium sulphate? Write the chemical equation involved in this process.**

**Q-6 Balance the following chemical equations-**

**(i) Mg(OH)2 + HCl → MgCl2 + H2O**

**(ii) N2 + H2 → NH3**

**(iii) Na + H2O → NaOH + H2**

**(iv) BaCl2 + Al2(SO4)3 → BaCl2 + AlCl3**

**Q-7 Write the balance chemical equations with state symbols for the following reactions-**

**(i) Iron fillings react with water to produce iron (III) oxide and hydrogen gas.**

**(ii) Sodium hydroxide solution reacts with aqueous solution of hydrochloric acid to form aqueous solution of sodium chloride and water.**

**Q-8 (i) State the physical condition of reaction in which any reaction will not take place.**

**(ii) What happens when iron pans or nails are kept in humid temperature?**

 **(iii) Why silver halides always kept in a dark color bottle?**

**Q-9 (i) Why food particles preferred to be packed in aluminium foil?**

 **(ii) Name two antioxidants which are generally used to prevent rancidity of food.**

**Q-10 Zinc oxide reacts with carbon, on heating to zinc metal and carbon mono oxide.**

 **(i) Write the balance chemical equation.**

 **(ii) Which are the oxidizing & reducing agent in the above reaction.**